



QUIZZES

Unit-6 (Chemical Bonding)



50 Questions



50 min

Topics

Unit-6 (Chemical Bonding)

Start Quiz

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49 : 58

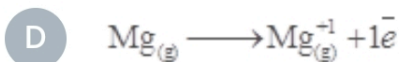
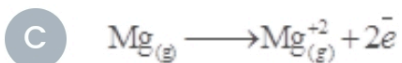
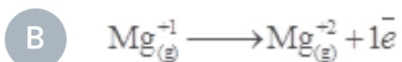
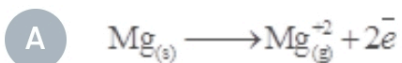


Q 1/50

50 min

i Hint

Q : Which of the following correctly represents the 2nd ionization energy of Mg



1

2

3

4

5

6

7

8

9

49 : 56



2/50



50 min



Hint

Q : General electronic configuration of p-block elements are given, Which of the following element has maximum first ionization energy



ns^2, np^2



ns^2, np^3



ns^2, np^4



ns^2, np^1

1

2

3

4

5

6

7

8

9

49 : 54



Q 3/50

50 min

Hint

Q : The molecule which has strongest bond is

A F-F

B Cl-Cl

C Br-Br

D I-I

1

2

3

4

5

6

7

8

9

49 : 53



4/50



50 min



Hint

Q : In which of the following compound, all carbon atoms show sp^2 hybridization



Carbon dioxide



Propene



1, 3-Butadiene



Ethane nitrile

1

2

3

4

5

6

7

8

9

49 : 51



Q 5/50

50 min

Hint

Q : Which of the following is most ionic in nature

A Sodium chloride

B Sodium fluoride

C Sodium bromide

D Sodium iodide

1

2

3

4

5

6

7

8

9

49 : 49



Q 6/50

50 min

Hint

Q : The compound which has maximum electronic repulsions is

A H_2O

B NH_3

C NH_4^+

D CH_4

1

2

3

4

5

6

7

8

9

49 : 48



7/50



50 min



Hint

Q : Octet rule is followed by the central atom of _____ molecule



BF₃



SO₂



SF₆



CCl₄

1

2

3

4

5

6

7

8

9

49 : 45



8/50



50 min



Hint

Q : Benzene contains delocalized π -electronic cloud due to



sp^2 - sp^2 overlapping



s-p overlapping



sp^2 -s overlapping



$p_z - p_z$ overlapping

1

2

3

4

5

6

7

8

9

49 : 44



9/50



50 min



Hint

Q :

Some properties are given below. Which corresponds to PCl_3

- (i) sp^2 hybridization (ii) Polar (iii) Trigonal pyramidal
(iv) AB_3 type molecule

A

i, ii, and iii

B

ii, iii and iv

C

ii and iii

D

i, ii, iii and iv

1

2

3

4

5

6

7

8

9

49 : 41



Q 10/50

50 min

Hint

Q : When bond order increases, then bond length becomes ____ and bond becomes ____

A Shorter, stronger

B Longer, weaker

C Shorter, weaker

D Longer, stronger

8

9

10

11

12

13

14

15

16

49 : 39



11/50



50 min



Hint

Q : The pair of compounds having same geometry and hybridization



SO₃, NH₃



H₂S, H₂O



CdCl₂, PCl₃



SO₂, BeCl₂

8

9

10

11

12

13

14

15

16

49 : 37



12/50



50 min



Hint

Q : Possible bonds formed by overlapping of p-p orbitals is/are



σ bond



Both σ and π



π -bond



None of these

8

9

10

11

12

13

14

15

16

49 : 36



Q 13/50

50 min

Hint

Q : % age of covalent bond in H_3O^+ is

A 33%

B 25%

C 75%

D 66%

8

9

10

11

12

13

14

15

16

49 : 35



Q 14/50

50 min

Hint

Q : Strength of bond depends upon the following factors except

A E.N difference between bonded atoms

B Sizes of the atoms

C Bond length

D Shielding effect

8

9

10

11

12

13

14

15

16

49 : 33



15/50



50 min



Hint

Q : The distance between the nuclei of two atoms forming covalent bond is called



Covalent radius



Bond angle



Atomic radius



Bond length

8

9

10

11

12

13

14

15

16

49 : 32



16/50



50 min



Hint

Q : Which one of the following has minimum electron affinity value



O



Se



S



Te

8

9

10

11

12

13

14

15

16

49 : 29



Q 17/50

50 min

Hint

Q : Which among the following has net dipole moment

A CH_4

B BF_3

C NH_3

D CCl_4

15

16

17

18

19

20

21

22

23

49 : 28



18/50



50 min



Hint

Q : Which overlapping may not lead to sigma bond formation



p-p in fluorine



s-p in hydrogen fluoride



sp^2-sp^2 in benzene



p-p in ethene

15

16

17

18

19

20

21

22

23

49 : 26



19/50



50 min



Hint

Q : Which of the following molecule contains maximum number of lone pairs



Chlorine



Oxygen



Carbon dioxide



Hydrogen chloride

15

16

17

18

19

20

21

22

23

49 : 25



20/50



50 min



Hint

Q : Correct statement when coordinate covalent bond is formed between NH_3 and BF_3



Ammonia is Lewis acid



Fluorine accepts lone pair due to its high electronegativity



Nitrogen of ammonia donates its lone pair to 2p orbital of Boron



Coordinate covalent bond is also called non-polar bond

15

16

17

18

19

20

21

22

23

49 : 24



Q 21/50

50 min

Hint

Q : Which one is AB_4 type molecule

A SO_3

B BF_3

C SO_2

D H_2S

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15

16

17

18

19

20

21

22

23

49 : 22



22/50



50 min



Hint

Q : According to VSEPR theory, which is angular in its structure



BF₃



SO₂



BeCl₂



NH₃

15

16

17

18

19

20

21

22

23

49 : 20



23/50



50 min



Hint

Q : During formation of ammonium ion from ammonia and hydrogen ion, there is no change in hybridization (sp^3) but bond angle is changed from 107.5° in ammonia to almost _____ in ammonium ion

A

104.5°

B

120°

C

109.5°

D

92°

20

21

22

23

24

25

26

27

28

49 : 18



24/50



50 min



Hint

Q :

In which of the following compound, central element is electron deficient



CH_4



BF_3



NH_3



SiCl_4

20

21

22

23

24

25

26

27

28

49 : 16



25/50



50 min



Hint

Q : Which one of the following shows minimum bond angle



BF_3



NF_3



CH_4



H_2O

20

21

22

23

24

25

26

27

28

49 : 15



26/50



50 min



Hint

Q : Which of following pair contains iso-structural species



CH_4 and CH_3^+



SO_4^{2-} and BF_4^-



NH_3 , BF_3



CO_2 and SO_2

20

21

22

23

24

25

26

27

28

49 : 13



27/50



50 min



Hint

Q : Which of the following compounds is non-polar

A

CHCl_3

B

CO

C

SO_2

D

CO_2

20

21

22

23

24

25

26

27

28

49 : 11



28/50



50 min



Hint

Q :

When the two partially filled atomic orbitals overlap in such a way that probability of finding the electron is maximum around the line joining the two nuclei, the result is the formation of

A

Pi bond

B

Sigma bond

C

Hydrogen bond

D

Metallic bond

20

21

22

23

24

25

26

27

28

49 : 07



29/50



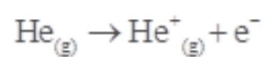
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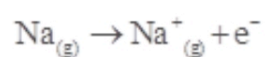
Hint

Q : Which one of the following processes requires the highest amount of energy

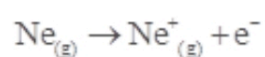
A



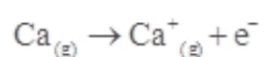
B



C



D



27

28

29

30

31

32

33

34

35

49 : 05



30/50



50 min



Hint

Q : The valence shell is



The highest energy level occupied by electrons



The set of orbitals used to make triple bonds



The orbitals belonging to the entire molecule



The lowest energy level occupied by electrons

27

28

29

30

31

32

33

34

35

49 : 04



31/50



50 min



Hint

Q : CH_4 is a nonpolar molecule. Which of the following similar molecules is also non-polar

A

CH_3Cl

B

SiH_3Cl

C

CH_2Cl_2

D

SiH_4

27

28

29

30

31

32

33

34

35

49 : 02



32/50



50 min



Hint

Q : The I.E. of nitrogen is more than that of oxygen due to the



The smallest size of nitrogen



The extra stability of the half-filled p orbitals



More penetrating effect



The greater attraction of electron by the nucleus

27

28

29

30

31

32

33

34

35

49 : 01



33/50



50 min



Hint

Q : The elements of group _____ shows abnormal low values of electron affinity in every period of periodic table



IIIA and VIA



IIA, VA and VIIA



IIA and VIIIA



IIA, VA and VIIIA

27

28

29

30

31

32

33

34

35

48 : 59



34/50



50 min



Hint

Q : The incorrect statement among the following is



The first ionization energy of Al is less than the first I.E. of Mg



The second ionization energy of Mg is greater than the second I.E. of Na



The first ionization energy of Na is less than the first I.E. of Mg



The third I.E. of Mg is greater than that of Al

27

28

29

30

31

32

33

34

35

48 : 57



Q 35/50

50 min

Hint

Q : Which one of the following will have smallest radius?

A Al^{3+}

B Mg^{2+}

C Si^{4+}

D Na^{2+}

27

28

29

30

31

32

33

34

35

48 : 55



36/50



50 min



Hint

Q : Which one of the following molecules has the highest dipole moment



H₂S



H₂O



SO₂



CS₂

33

34

35

36

37

38

39

40

41

48 : 53



37/50



50 min



Hint

Q : All the atoms are coplanar in the molecule_____



CH₄



PH₃



BF₃



NH₃

33

34

35

36

37

38

39

40

41

48 : 52



Q 38/50

50 min

Hint

Q : Bonding in phosphonium ion is _____ percent covalent

A 25

B 50

C 33

D 75

33

34

35

36

37

38

39

40

41

48 : 49



39/50



50 min



Hint

Q : Formation of MgO is an example of



Ionic bond



Polar covalent bond



Non-polar Covalent bond



Double Covalent bond

33

34

35

36

37

38

39

40

41

48 : 48



40/50



50 min



Hint

Q : The bond between H-H is



Stronger than the bond between C-C



Weaker than the bond between C-C



Neither stronger nor weaker than the bond between C-C



Not comparable

33

34

35

36

37

38

39

40

41

48 : 46



41/50



50 min



Hint

Q : Which hybrid orbitals are used for bonding in triangular pyramidal molecule



sp^2



dsp^2



sp



sp^3

33

34

35

36

37

38

39

40

41

48 : 44



42/50



50 min



Hint

Q : The hybridization state of 'S' in SO_3 is similar to that of



C in C_2H_2



C in CO_2



C in C_2H_6



C in C_2H_4

40

41

42

43

44

45

46

47

48

48 : 42



Q 43/50

50 min

Hint

Q : Which of the following is NOT tetrahedral

A BF_4^-

B SO_4^{2-}

C NH_4^+

D CO_3^{2-}

40

41

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43

44

45

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47

48

48 : 41



44/50

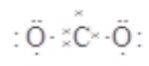
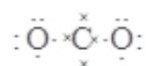
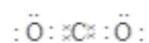
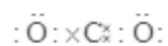


50 min



Hint

Q : Which one is correct dot and cross diagram of CO₂



40

41

42

43

44

45

46

47

48

48 : 40



45/50



50 min



Hint

Q : The electron affinity is the measure of the attraction of the nucleus of an atom for



Valance electron



Inner shell electron



Extra incoming electron



Last electron

40

41

42

43

44

45

46

47

48

48 : 38



46/50



50 min



Hint

Q : The ionization energies of an element are given below



The element X may belong to

A

IVA

B

IIIA

C

IIA

D

IA

40

41

42

43

44

45

46

47

48

48 : 37



47/50



50 min



Hint

Q : In the resonance structure of benzene the number of s-bonds and p delocalized electrons are respectively



12 and 6



6 and 6



6 and 3



12 and 3

40

41

42

43

44

45

46

47

48

48 : 35



48/50



50 min



Hint

Q : The hybridization associated with the central atom of a molecule in which all the bond angles are 120° is



sp



sp³



sp²



dsp³

40

41

42

43

44

45

46

47

48

48 : 33



49/50



50 min



Hint

Q :

Choose the species that is incorrectly matched with the shape of the central atom



CF₄ tetrahedral



H₂O tetrahedral



BeCl₂ linear



NH₃ pyramidal

42

43

44

45

46

47

48

49

50

48 : 31



50/50



50 min



Hint

Q : Correct order of H-N-H angle for the species NH_3 , NH_4^+ and NH_2^-

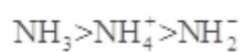
A



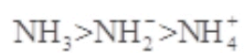
B



C



D



42

43

44

45

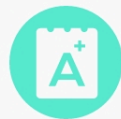
46

47

48

49

50



QUIZ RESULT

Unit-6 (Chemical Bonding)



50



50 min



15-Jun-2021



0 sec



0/50



0.0%

Result Detail



| | |
|-------------|----|
| Correct | 0 |
| Incorrect | 0 |
| Unattempted | 50 |



Chemistry

0%



Unit-6 (Chemical Bonding)



Correct



Unattempted

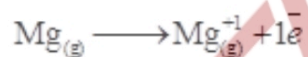
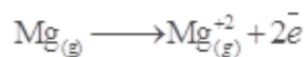
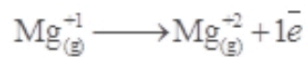
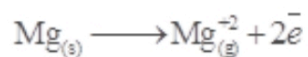


Incorrect



1/50

Q : Which of the following correctly represents the 2nd ionization energy of Mg



1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



2/50

Q : General electronic configuration of p-block elements are given, Which of the following element has maximum first ionization energy



ns^2, np^2



ns^2, np^3



ns^2, np^4



ns^2, np^1

1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



3/50

Q : The molecule which has strongest bond is

A

F-F

B

Cl-Cl

C

Br-Br

D

I-I

1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



4/50

Q : In which of the following compound, all carbon atoms show sp^2 hybridization

A

Carbon dioxide

B

Propene

C

1, 3-Butadiene

D

Ethane nitrile

1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



5/50

Q : Which of the following is most ionic in nature

A

Sodium chloride

B

Sodium fluoride

C

Sodium bromide

D

Sodium iodide

1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



6/50

Q : The compound which has maximum electronic repulsions is

A



B



C



D



1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



7/50

Q : Octet rule is followed by the central atom of _____ molecule

A

BF_3

B

SO_2

C

SF_6

D

CCl_4

1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



8/50

Q : Benzene contains delocalized π -electronic cloud due to

A

sp^2 - sp^2 overlapping

B

s-p overlapping

C

sp^2 -s overlapping

D

$p_z - p_z$ overlapping

1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



9/50

Q :

Some properties are given below. Which corresponds to PCl_3

- (i) sp^2 hybridization (ii) Polar (iii) Trigonal pyramidal
(iv) AB_3 type molecule

A

i, ii, and iii

B

ii, iii and iv

C

ii and iii

D

i, ii, iii and iv

1

2

3

4

5

6

7

8

9



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



10/50

Q : When bond order increases, then bond length becomes _____ and bond becomes ____

A

Shorter, stronger

B

Longer, weaker

C

Shorter, weaker

D

Longer, stronger



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



11/50

Q : The pair of compounds having same geometry and hybridization

A

SO_3, NH_3

B

$\text{H}_2\text{S}, \text{H}_2\text{O}$

C

$\text{CdCl}_2, \text{PCl}_3$

D

$\text{SO}_2, \text{BeCl}_2$



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



12/50

Q : Possible bonds formed by overlapping of p-p orbitals is/are

A

σ bond

B

Both σ and π

C

π -bond

D

None of these

9

10

11

12

13

14

15

16

17



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



13/50

Q : % age of covalent bond in H_3O^+ is

A

B

C

D

9

10

11

12

13

14

15

16

17



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



14/50

Q : Strength of bond depends upon the following factors except

A

E.N difference between bonded atoms

B

Sizes of the atoms

C

Bond length

D

Shielding effect

9

10

11

12

13

14

15

16

17



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



15/50

Q : The distance between the nuclei of two atoms forming covalent bond is called

A

Covalent radius

B

Bond angle

C

Atomic radius

D

Bond length



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



16/50

Q : Which one of the following has minimum electron affinity value

A

O

B

Se

C

S

D

Te

9

10

11

12

13

14

15

16

17



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



17/50

Q : Which among the following has net dipole moment

A

CH_4

B

BF_3

C

NH_3

D

CCl_4



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



18/50

Q : Which overlapping may not lead to sigma bond formation

A

p-p in fluorine

B

s-p in hydrogen fluoride

C

sp^2-sp^2 in benzene

D

p-p in ethene



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



19/50

Q : Which of the following molecule contains maximum number of lone pairs

A

Chlorine

B

Oxygen

C

Carbon dioxide

D

Hydrogen chloride



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



20/50

Q : Correct statement when coordinate covalent bond is formed between NH_3 and BF_3



Ammonia is Lewis acid



Fluorine accepts lone pair due to its high electronegativity



Nitrogen of ammonia donates its lone pair to 2p orbital of Boron



Coordinate covalent bond is also called non-polar bond



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



21/50

Q : Which one is AB_4 type molecule

A

SO_3

B

BF_3

C

SO_2

D

H_2S



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



22/50

Q : According to VSEPR theory, which is angular in its structure

A

BF_3

B

SO_2

C

BeCl_2

D

NH_3



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



23/50

Q : During formation of ammonium ion from ammonia and hydrogen ion, there is no change in hybridization (sp^3) but bond angle is changed from 107.5° in ammonia to almost _____ in ammonium ion

A

104.5°

B

120°

C

109.5°

D

92°



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



24/50

Q :

In which of the following compound, central element is electron deficient

A



B



C



D





Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



25/50

Q : Which one of the following shows minimum bond angle

A



B



C



D





Unit-6 (Chemical Bonding)



Correct



Unattempted



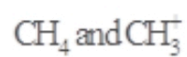
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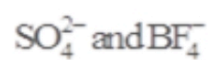
26/50

Q : Which of following pair contains iso-structural species

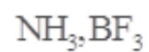
A



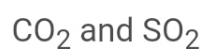
B



C



D





Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



27/50

Q : Which of the following compounds is non-polar

A

CHCl_3

B

CO

C

SO_2

D

CO_2



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



28/50

Q :

When the two partially filled atomic orbitals overlap in such a way that probability of finding the electron is maximum around the line joining the two nuclei, the result is the formation of

A

Pi bond

B

Sigma bond

C

Hydrogen bond

D

Metallic bond



Unit-6 (Chemical Bonding)



Correct



Unattempted



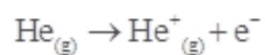
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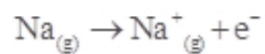
29/50

Q : Which one of the following processes requires the highest amount of energy

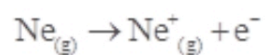
A



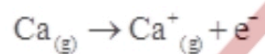
B



C



D





Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



30/50

Q : The valence shell is

A

The highest energy level occupied by electrons

B

The set of orbitals used to make triple bonds

C

The orbitals belonging to the entire molecule

D

The lowest energy level occupied by electrons



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



31/50

Q : CH_4 is a nonpolar molecule. Which of the following similar molecules is also non-polar

A

CH_3Cl

B

SiH_3Cl

C

CH_2Cl_2

D

SiH_4



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



32/50

Q : The I.E. of nitrogen is more than that of oxygen due to the

A

The smallest size of nitrogen

B

The extra stability of the half-filled p orbitals

C

More penetrating effect

D

The greater attraction of electron by the nucleus



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



33/50

Q : The elements of group _____ shows abnormal low values of electron affinity in every period of periodic table

A

IIIA and VIA

B

IIA, VA and VIIA

C

IIA and VIIIA

D

IIA, VA and VIIIA



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



34/50

Q : The incorrect statement among the following is



The first ionization energy of Al is less than the first I.E. of Mg



The second ionization energy of Mg is greater than the second I.E. of Na



The first ionization energy of Na is less than the first I.E. of Mg



The third I.E. of Mg is greater than that of Al



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



35/50

Q : Which one of the following will have smallest radius?

A

Al^{3+}

B

Mg^{2+}

C

Si^{4+}

D

Na^{2+}



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



36/50

Q : Which one of the following molecules has the highest dipole moment

A



B



C



D





Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



37/50

Q : All the atoms are coplanar in the molecule_____

A

CH_4

B

PH_3

C

BF_3

D

NH_3



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



38/50

Q : Bonding in phosphonium ion is _____ percent covalent

A

25

B

50

C

33

D

75

34

35

36

37

38

39

40

41

42



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



39/50

Q : Formation of MgO is an example of

A

Ionic bond

B

Polar covalent bond

C

Non-polar Covalent bond

D

Double Covalent bond

34

35

36

37

38

39

40

41

42



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



40/50

Q : The bond between H-H is

A

Stronger than the bond between C-C

B

Weaker than the bond between C-C

C

Neither stronger nor weaker than the bond between C-C

D

Not comparable



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



41/50

Q : Which hybrid orbitals are used for bonding in triangular pyramidal molecule

A

sp^2

B

dsp^2

C

sp

D

sp^3



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



42/50

Q : The hybridization state of 'S' in SO_3 is similar to that of

A

C in C_2H_2

B

C in CO_2

C

C in C_2H_6

D

C in C_2H_4



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



43/50

Q : Which of the following is NOT tetrahedral

A



B



C



D





Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



44/50

Q : Which one is correct dot and cross diagram of CO_2

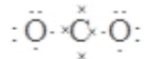
A



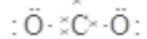
B



C



D





Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



45/50

Q : The electron affinity is the measure of the attraction of the nucleus of an atom for

A

Valance electron

B

Inner shell electron

C

Extra incoming electron

D

Last electron



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



46/50

Q : The ionization energies of an element are given below



The element X may belong to

A

IVA

B

IIIA

C

IIA

D

IA



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



47/50

Q : In the resonance structure of benzene the number of s-bonds and p delocalized electrons are respectively

A

12 and 6

B

6 and 6

C

6 and 3

D

12 and 3



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



48/50

Q : The hybridization associated with the central atom of a molecule in which all the bond angles are 120° is

A

sp

B

sp³

C

sp²

D

dsp³



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



49/50

Q :

Choose the species that is incorrectly matched with the shape of the central atom

A

CF_4 tetrahedral

B

H_2O tetrahedral

C

BeCl_2 linear

D

NH_3 pyramidal



Unit-6 (Chemical Bonding)



Correct



Unattempted



Incorrect



50/50

Q : Correct order of H-N-H angle for the species NH_3 , NH_4^+ and NH_2^-

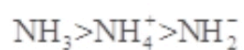
A



B



C



D

